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8. Explain the difference between ionic and covalent bonds.
ionic \rightarrow e⁻ are lost or gained in ions and opp charge
covalent \rightarrow e⁻ are shared

9. For the following chemical equations, balance and classify.
a) $\text{Ca(OH)}_2 + 2\text{HCl} \rightarrow \text{CaCl}_2 + 2\text{H}_2\text{O}$ double displacement
b) $2\text{Ag}_2\text{O} \rightarrow 4\text{Ag} + \text{O}_2$ decomposition

10. Write the chemical formulas for the following:
a) lead (II) nitrate $\text{Pb(NO}_3)_2$
b) silver acetate $\text{AgC}_2\text{H}_3\text{O}_2$
c) potassium oxide K_2O
d) zinc phosphate $\text{Zn}_3(\text{PO}_4)_2$
e) magnesium chloride MgCl_2

11. Write the name for the following:
a) CuCl_2 copper(II) chloride
b) LiBr lithium bromide
c) HNO_3 nitric acid
d) Na_2O sodium oxide

12. For the following, predict the products (don't forget the subscripts) and balance the equation.
 $2\text{HCl(aq)} + \text{PbCl}_2(\text{s}) \rightarrow \text{PbCl}_2(\text{s}) + 2\text{HCl(aq)}$

13. For question 12 write a total ionic equation and a net ionic equation.
 $2\text{H}^+(\text{aq}) + 2\text{Cl}^-(\text{aq}) + \text{Pb}^{2+}(\text{aq}) + 2\text{Cl}^-(\text{aq}) \rightarrow \text{PbCl}_2(\text{s}) + 2\text{H}^+(\text{aq}) + 2\text{Cl}^-(\text{aq})$
Net ionic: $\text{Pb}^{2+} + 2\text{Cl}^- \rightarrow \text{PbCl}_2(\text{s})$

14. Use your solubility table to determine whether the substance will form a precipitate or not.
a) magnesium sulfide no precip
b) calcium carbonate precipitate
c) PbCl_2 precipitate
d) $\text{Ba}_3(\text{PO}_4)_2$ no precip

15. An unknown solution may contain Fe^{3+} and/or Sr^{2+} . Use a flowchart to present how you would find out if one or both are present in the solution. (Use a separate sheet of paper)

① add SO_4^{2-}
if precip then Sr present
if no precip then no Sr present
 $\text{SrSO}_4(\text{s})$

② add S^{2-} or OH^- test
if precip then Fe^{3+} present
if no precip then Fe^{3+} not present
 $\text{Fe}(\text{OH})_3$ or Fe_2S_3

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