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Finally I get this ebook, thanks for all these I can get now!

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Cool! I'am really happy

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My friends are so mad that they do not know how I have all the high quality ebook which they do not!

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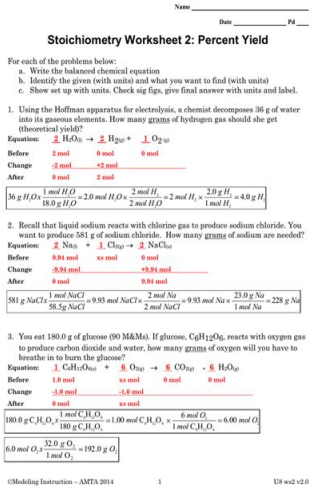
so many fake sites. this is the first one which worked! Many thanks



Name _____

Date _____ Pd ___

Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:
a. Write the balanced chemical equation
b. Identify the given (with units) and what you want to find (with units)
c. Show set up with units. Check sig figs, give final answer with units and label.

1. Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?

Equation: 2 $H_2O_{(l)}$ → 2 $H_{2(g)}$ + 1 $O_{2\ (g)}$

Before	2 mol	0 mol	0 mol
Change	-2 mol	+2 mol	
After	0 mol	2 mol	

$$36\ g\ H_2O \times \frac{1\ mol\ H_2O}{18.0\ g\ H_2O} = 2.0\ mol\ H_2O \times \frac{2\ mol\ H_2}{2\ mol\ H_2O} = 2\ mol\ H_2 \times \frac{2.0\ g\ H_2}{1\ mol\ H_2} = 4.0\ g\ H_2$$

2. Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?

Equation: 2 $Na_{(l)}$ + 1 $Cl_{2(g)}$ → 2 $NaCl_{(s)}$

Before	9.94 mol	xs mol	0 mol
Change	-9.94 mol		+9.94 mol
After	0 mol		9.94 mol

$$581\ g\ NaCl \times \frac{1\ mol\ NaCl}{58.5\ g\ NaCl} = 9.93\ mol\ NaCl \times \frac{2\ mol\ Na}{2\ mol\ NaCl} = 9.93\ mol\ Na \times \frac{23.0\ g\ Na}{1\ mol\ Na} = 228\ g\ Na$$

3. You eat 180.0 g of glucose (90 M&Ms). If glucose, $C_6H_{12}O_6$, reacts with oxygen gas to produce carbon dioxide and water, how many grams of oxygen will you have to breathe in to burn the glucose?

Equation: 1 $C_6H_{12}O_{6(s)}$ + 6 $O_{2(g)}$ → 6 $CO_{2(g)}$ + 6 $H_2O_{(g)}$

Before	1.0 mol	xs mol	0 mol	0 mol
Change	-1.0 mol	-1.0 mol		
After	0 mol	xs mol		

$$180.0\ g\ C_6H_{12}O_6 \times \frac{1\ mol\ C_6H_{12}O_6}{180\ g\ C_6H_{12}O_6} = 1.00\ mol\ C_6H_{12}O_6 \times \frac{6\ mol\ O_2}{1\ mol\ C_6H_{12}O_6} = 6.00\ mol\ O_2$$

$$6.0\ mol\ O_2 \times \frac{32.0\ g\ O_2}{1\ mol\ O_2} = 192.0\ g\ O_2$$

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